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Answers: Moles and Stoichiometry Practice Problems 1) How many moles of sodium atoms correspond to 1.56×10^{21} atoms of sodium? 1.56×10^{21} atoms Na $\times 1 \text{ mol Na} = 2.59 \times 10^3 \text{ mol Na}$ $236.022 \times 10 \text{ atoms Na}$ 2) Determine the mass in grams of each of the following: a. 1.35 mol of Fe $1.35 \text{ mol Fe} \times 55.845 \text{ g Fe} = 75.4 \text{ g Fe}$ 1 mol Fe b. 24.5 mol O

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Chapter 9 Stoichiometry Test Answers

SOLUTION: You are asked to calculate the amount (in moles) of oxygen atoms in a given sample. You are given the mass of the oxygen sample. Use the molar mass of oxygen ($1 \text{ mol O} = 16.00 \text{ g}$) to create a conversion factor that converts mass (in grams) to amount (in mol) of oxygen. $124 \text{ g O} \times 1 \text{ mol O} / 16.00 \text{ g} = 7.75 \text{ mol O}$.

Chapter 3 Stoichiometry

It is readily available in pdf, ppt, word, rar, txt, kindle, and also zip. Answer Key Mole/Stoichiometry.Test.Review. Answer Key

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Mole/Stoichiometry.Test.Review 1. 6.022×10^{23} particles ((atoms, (molecules) ((2. 1 mole (= 6.022×10^{23} particles ((1 mole=molar (mass (1 mole=22.4L (3.

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