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which the starting materials (reactants) are different from the products. Chemical reactions tend to involve the motion of electrons, leading to the formation and breaking of chemical bonds. There are several

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substance B we must use moles. $A+B$

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products A + other reactant B A & B can
be any two reactants or products Mass of
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one or more substances, also called

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reactants, are converted to one or more different substances, known as products. Substances are either chemical elements or compounds. A chemical reaction rearranges the constituent atoms of the reactants to create different substances as products. The properties of the products are different from those of the reactants.

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Technetium can, however, be prepared by nuclear transmutation reactions. For example, bombarding a molybdenum-96 target with deuterium nuclei $(\text{}_{1}^{2}\text{H})$ produces technetium-97. Identify the other product of the reaction and write a balanced nuclear equation for this

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textbook/reference ever to cover the chemical basis of life, the "Green Bible of Biochemistry" has been a well-respected contribution to the field for more than twenty years. The complex structures that make up cells are described in detail, along with the forces that hold them together, and the chemical reactions that

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actively pursuing the molecular basis of thermo dynamics, work that culminated in the Boltzmann equation and the theory of transport processes in dilute gases. Much later, Onsager undertook the elucidation of the symmetry of transport coefficients and, thereby, established himself as the father of the theory of nonequilibrium

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thermodynamics. Combining the statistical ideas of Gibbs and Langevin with the phenomenological transport equations, Onsager and others went on to develop a consistent statistical theory of irreversible processes. The power of that theory is in its ability to relate measurable quantities, such as transport coefficients

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and thermodynamic derivatives, to the results of experimental measurements. As powerful as that theory is, it is linear and limited in validity to a neighborhood of equilibrium. In recent years it has been possible to extend the statistical theory of nonequilibrium processes to include nonlinear effects. The modern theory, as

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expounded in this book, is applicable to a wide variety of systems both close to and far from equilibrium. The theory is based on the notion of elementary molecular processes, which manifest themselves as random changes in the extensive variables characterizing a system. The theory has a hierarchical character and, thus, can be

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not incorporated into detailed chemical kinetic models, and the use of modern computational quantum chemistry, which has recently become an extraordinarily useful component of the reaction kinetics toolkit. It is intended both for those who need to model complex chemical reaction processes but have little background in the

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